Received 12 July 2006

Accepted 15 September 2006

Acta Crystallographica Section E Structure Reports Online

ISSN 1600-5368

## Fang-Fang Jian\* and Ying Li

New Materials & Function Coordination Chemistry Laboratory, Qingdao University of Science & Technology, Qingdao 266042, People's Republic of China

Correspondence e-mail: ffj2003@163169.net

#### Key indicators

Single-crystal X-ray study T = 293 KMean  $\sigma(\text{C}-\text{C}) = 0.003 \text{ Å}$  R factor = 0.039 wR factor = 0.111 Data-to-parameter ratio = 17.1

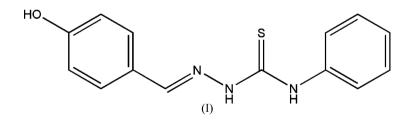
For details of how these key indicators were automatically derived from the article, see http://journals.iucr.org/e.

# 1-(4-Hydroxybenzylidene)-4-phenylthiosemicarbazide

The title compound,  $C_{14}H_{13}N_3OS$ , was prepared by the reaction of *p*-hydroxybenzaldehyde, hydrazine and phenyl isothiocyanate. The crystal packing is stabilized by  $C-H\cdots\pi$ ,  $N-H\cdots S$ ,  $C-H\cdots O$  and  $O-H\cdots S$  intermolecular interactions.

### Comment

Thiourea (TU) is a very convenient and frequently employed nucleophile used to study ligand substitution reactions in coordination chemistry because of its good solubility, neutral character and high nucleophilicity (Schiessl *et al.*, 2005). Thiourea derivatives have been successfully screened for various biological activities (Antholine & Taketa, 1982). Furthermore, TU and its derivatives have also been screened for allergenic and carcinogenic factors. It has also been shown that their presence inhibits nitrification in soil and water (Spataru *et al.*, 2005). As TU and its derivatives have unique characteristics and pharmacutical potential, we have synthesized the title compound, (I), and describe its structure here.



In the title compound, bond lengths and angles are in usual ranges (Ji *et al.*, 2002). In (I), atoms S1, N1, N2, C6 and C7 are planar (plane p1); the intramolecular hydrogen bond N1—H1B···N3 (Table 1 and Fig. 1) contributes to the planarity. The other four atoms (N2, N3, C8, C9) are also coplanar (plane p2). The dihedral angle between the two aromatic rings is 83.51 (3)°. The dihedral angle between p1 and p2 is 12.30 (2)°.

The crystal packing is realised by O-H···S and N-H···S hydrogen bonds (Table 1 and Fig. 2). Molecules are involved in C-H··· $\pi$  interactions between C13-H13 and the centroid (*Cg*) of the C1-C6 ring at (1 + *x*,  $\frac{1}{2} - y$ ,  $-\frac{1}{2} + z$ ); C-H = 0.93 Å, H···*Cg* = 2.80 Å, C···*Cg* = 3.595 (2) Å and C-H···*Cg* = 144°.

## **Experimental**

The title compound was prepared by the reaction of hydrazine (1.0 g, 20 mmol) and *p*-hydroxybenzaldehyde (2.4 g, 20 mmol) with phenyl isothiocyanate (2.7 g, 20 mmol). Single crystals of the title compound suitable for X-ray measurements were obtained by recrystallization

Acta Cryst. (2006). E**62**, 04563–04564

All rights reserved

© 2006 International Union of Crystallography

# organic papers

from an ethanol solution at room temperature. (yield 5.0 g, 92.3%; m.p. 481-484 K).

Z = 4

 $D_x = 1.304 \text{ Mg m}^{-3}$ 

1914 reflections with  $I > 2\sigma(I)$ 

3 standard reflections

every 100 reflections

intensity decay: none

 $w = 1/[\sigma^2(F_0^2) + (0.0455P)^2]$ + 0.3594P]

 $(\Delta/\sigma)_{\rm max} < 0.001$  $\Delta \rho_{\text{max}} = 0.27 \text{ e } \text{\AA}^{-3}$  $\Delta \rho_{\text{min}} = -0.18 \text{ e } \text{\AA}^{-3}$ 

where  $P = (F_0^2 + 2F_c^2)/3$ 

Extinction correction: SHELXL97 Extinction coefficient: 0.0094 (15)

Mo  $K\alpha$  radiation

 $\mu = 0.23 \text{ mm}^{-1}$ 

T = 293 (2) K

Block, yellow  $0.30 \times 0.25 \times 0.25$  mm

 $R_{int} = 0.013$  $\theta_{\rm max} = 27.0^{\circ}$ 

### Crystal data

C14H13N3OS  $M_r = 271.34$ Monoclinic,  $P2_1/c$ a = 10.323 (2) Å b = 12.697 (3) Å c = 10.597 (2) Å  $\beta = 95.57 (3)^{\circ}$ V = 1382.4 (5) Å<sup>3</sup>

#### Data collection

Enraf-Nonius CAD-4 diffractometer  $\omega$  scans Absorption correction: none 3123 measured reflections 2999 independent reflections

#### Refinement

Refinement on  $F^2$  $R[F^2 > 2\sigma(F^2)] = 0.039$  $wR(F^2) = 0.111$ S = 1.022964 reflections 173 parameters H-atom parameters constrained

Table 1

Hydrogen-bond geometry (Å, °).

$D - H \cdot \cdot \cdot A$	D-H	$H \cdot \cdot \cdot A$	$D \cdots A$	$D - \mathbf{H} \cdots \mathbf{A}$
$O1-H1A\cdots S1^{i}$ $N1-H1B\cdots N3$	0.82	2.43 2.18	3.159 (3)	149
$N1 - H1B \cdots N3$ $N2 - H2A \cdots S1^{ii}$	0.86 0.86	2.18 2.62	2.594 (2) 3.422 (2)	109 156

Symmetry codes: (i) x + 1,  $-y + \frac{1}{2}$ ,  $z - \frac{1}{2}$ ; (ii) -x + 2, -y, -z.

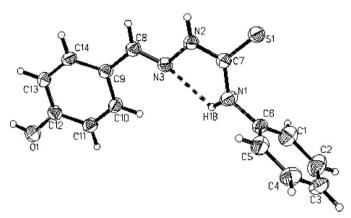
All H atoms were positioned geometrically (N-H = 0.86, C-H =0.93 Å) and refined as riding, with  $U_{iso}(H) = 1.2U_{eq}(C,N)$ .

Data collection: CAD-4 Software (Enraf-Nonius, 1989); cell refinement: CAD-4 Software; data reduction: NRCVAX (Gabe et al., 1989); program(s) used to solve structure: SHELXS97 (Sheldrick, 1997); program(s) used to refine structure: SHELXL97 (Sheldrick, 1997); molecular graphics: SHELXTL/PC (Sheldrick, 1990); software used to prepare material for publication: WinGX (Farrugia, 1999).

The authors thank the Natural Science Foundation of Shandong Province (No. Y2005B04).

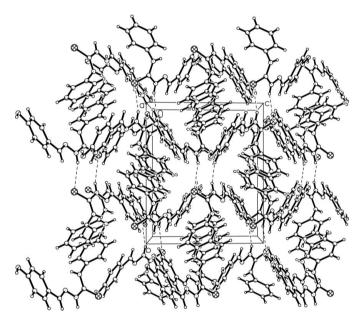
## References

Antholine, W. & Taketa, F. (1982). J. Inorg. Biochem. 16, 145-154. Enraf-Nonius (1989). CAD-4 Software. Version 5.0. Enraf-Nonius, Delft, The Netherlands.



#### Figure 1

The molecular structure of the title compound, showing 30% probability displacement ellipsoids and the atom-numbering scheme. The intramolecular hydrogen bond is indicated by a dashed line.



#### Figure 2

Crystal packing of (I). Hydrogen bonds are shown as dashed lines.

Farrugia, L. J. (1999). J. Appl. Cryst. 32, 837-838.

- Gabe, E. J., Le Page, Y., Charland, J.-P., Lee, F. L. & White, P. S. (1989). J. Appl. Cryst. 22, 384-387.
- Ji, B. M., Du, C. X., Zhu, Y. & Wang, Y. (2002). Chin. J. Struct. Chem. 21, 252-255.
- Schiessl, W. C., Summa, N. K., Weber, C. F., Gubo, S., Benfer, C. D., Puchta, R., Van Eikema Hommes, N. J. R. & Van Eldik, R. (2005). Z. Anorg. Allg. Chem. 631, 2812-2819.
- Sheldrick, G. M. (1990). SHELXTL/PC. Siemens Analytical X-ray Instruments Inc., Madison, Wisconsin, USA.
- Sheldrick, G. M. (1997). SHELXL97 and SHELXS97. University of Göttingen, Germany.
- Spataru, N., Spataru, T. & Fujishima, A. (2005). Electroanalysis, 17, 800-805.